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Key indicators

Single-crystal X-ray study T = 298 KMean σ (C–C) = 0.006 Å Disorder in main residue R factor = 0.061 wR factor = 0.200 Data-to-parameter ratio = 13.7

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

1,5-Diphenyl-3-(2-thienyl)pentane-1,5-dione

The title compound, $C_{21}H_{18}O_2S$, was synthesized by the reaction of thiophene-2-carbaldehyde with acetophenone and NaOH. All bond lengths and angles show normal values. The crystal packing is stabilized by weak intermolecular $C-H\cdots O$ hydrogen bonds and van der Waals forces.

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Comment

The drive towards clean technology has encouraged the application of solvent-free conditions in industrial chemistry (Dittmer, 1997). A move away from the use of solvents in organic synthesis has led in some cases to improved results and more benign synthetic procedures (Kidwai *et al.*, 2001). Adopting the principles of 'green chemistry', we have established that using solvent-free conditions for condensation and Michael reactions results in a dramatic improvement in yields. Our approach reduces the use of organic solvents, which are potentially toxic and hazardous (Anastas & Varner, 1998), and uses simple and mild conditions, with inherently lower costs (Tanaka & Toda, 2000). In view of the current emphasis on 'green chemistry', we set out to develop a solvent-free preparation of coumarins using an inexpensive and non-polluting catalyst.



The synthesis of substituted thiophenes has attracted a great deal of interest over the years due to their presence in



© 2006 International Union of Crystallography All rights reserved View of (I), showing the atomic labelling and displacement ellipsoids drawn at the 30% probability level. Only major components of the disordered S1 and C3 atoms are shown.

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natural products (Koike *et al.*, 1999). 1,5-Diketones are important synthetic intermediates and starting materials in the synthesis of many heterocyclic compounds (Hirsch & Bailey, 1978; Krohnke, 1976). In a continuation of previous work on the synthesis of 1,5-diketones (Constable *et al.*, 1998; Fuchigami *et al.*, 1986), we present a new title compound, (I), synthesized under solvent-free conditions.

In (I) (Fig. 1), the bond lengths and angles are normal and similar to those observed in 1,3,5-triphenylpentane-1,5-dione (Das *et al.*, 1994). Weak intermolecular $C-H \cdots O$ hydrogen bonds (Table 1) link the molecules into chains extending along the *c* axis. The crystal packing (Fig. 2) is further stabilized by van der Waals forces.

Experimental

Acetophenone (0.75 g, 6.25 mmol), freshly distilled thiophene-2carbaldehyde (0.35 g, 3.125 mmol) and NaOH (0.25 g, 6.25 mmol) were aggregated with a glass paddle in an open flask. The resulting mixture was washed with water several times to remove NaOH and was recrystallized from ethanol, affording the title compound as a crystalline solid.

Z = 4

 $D_x = 1.272 \text{ Mg m}^{-3}$

 $0.58 \times 0.55 \times 0.53~\text{mm}$

8511 measured reflections

 $R_{\rm int} = 0.060$

 $\theta_{\rm max} = 25.0^{\circ}$

3061 independent reflections 1620 reflections with $I > 2\sigma(I)$

Mo $K\alpha$ radiation $\mu = 0.20 \text{ mm}^{-1}$ T = 298 (2) KBlock, colourless

Crystal data

$C_{21}H_{18}O_2S$
$M_r = 334.41$
Monoclinic, $P2_1/c$
a = 13.460 (2) Å
b = 12.094 (3) Å
c = 10.762 (2) Å
$\beta = 94.456 \ (3)^{\circ}$
V = 1746.6 (6) Å ³

Data collection

Bruker SMART CCD area-detector
diffractometer
φ and ω scans
Absorption correction: multi-scan
(SADABS; Sheldrick, 1996)
$T_{\min} = 0.896, \ T_{\max} = 0.904$

Refinement

Refinement on F^2	$w = 1/[\sigma^2(F_0^2) + (0.0999P)^2$
$R[F^2 > 2\sigma(F^2)] = 0.061$	+ 0.3043P]
$wR(F^2) = 0.200$	where $P = (F_0^2 + 2F_c^2)/3$
S = 1.00	$(\Delta/\sigma)_{\rm max} < 0.001$
3061 reflections	$\Delta \rho_{\rm max} = 0.23 \ {\rm e} \ {\rm \AA}^{-3}$
224 parameters	$\Delta \rho_{\rm min} = -0.31 \text{ e } \text{\AA}^{-3}$
H-atom parameters constrained	

Table 1

Hydrogen-bond geometry (Å, °).

$D - H \cdot \cdot \cdot A$	D-H	$H \cdot \cdot \cdot A$	$D \cdots A$	$D - H \cdots A$
$C6-H6B\cdotsO1^{i}$	0.97	2.47	3.361 (7)	153
	2 1			

Symmetry code: (i) $x, -y + \frac{3}{2}, z - \frac{1}{2}$.

All H atoms were positioned geometrically and refined using a riding model, with C–H = 0.93–0.98 Å and $U_{\rm iso}({\rm H})$ = 1.2–1.5 $U_{\rm eq}({\rm C})$. Atoms S1, C3 and H3 were treated as disordered between two positions, with refined occupancies of 0.735 (5) and 0.265 (5).



Figure 2

The crystal packing of (I), viewed along the *b* axis. The dashed lines denote $C-H\cdots O$ hydrogen bonds. Only the major components of the disordered S1, C3 and H3 atoms are shown for each molecule.

Data collection: *SMART* (Siemens, 1996); cell refinement: *SAINT* (Siemens, 1996); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997*a*); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997*a*); molecular graphics: *SHELXTL* (Sheldrick, 1997*b*); software used to prepare material for publication: *SHELXTL*.

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References

- Anastas, P. T. & Varner, J. C. (1998). Green Chemistry, Theory and Practice, pp. 1–119. Oxford University Press.
- Constable, E. C., Neuburger, M., Smith, D. R. & Zehnder, M. (1998). Inorg. Chim. Acta, 275–276, 359–365.
- Das, G. C., Hursthouse, M. B., Malik, K. M. A., Rahman, M. M., Rahman, M. T. & Olsson, T. (1994). J. Chem. Crystallogr. 24, 511–515.
- Dittmer, D. C. (1997). Chem. Ind. pp. 779-784.
- Fuchigami, T., Awata, T., Nonaka, T. & Baizer, M. M. (1986). Bull. Chem. Soc. Jpn, 59, 2873–2879.
- Hirsch, S. S. & Bailey, W. J. (1978). J. Org. Chem. 43, 4090-4094.
- Kidwai, M., Sapra, P., Bhushan, K. R. & Misra, P. (2001). Synthesis, 10, 1509– 1512.
- Koike, K., Jia, Z., Nikaib, T., Liu, Y. & Guo, D. (1999). Org. Lett. 1, 197–198. Krohnke, F. (1976). Synthesis, 1, 1–24.
- Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.
- Sheldrick, G. M. (1997a). SHELXS97 and SHELXL97. University of Göttingen, Germany.
- Sheldrick, G. M. (1997b). SHELXTL. Version 5.1. Bruker AXS Inc., Madison, Wisconsin, USA.
- Siemens (1996). *SMART* and *SAINT*. Siemens Analytical X-ray Instruments Inc., Madison, Wisconsin, USA.
- Tanaka, K. & Toda, F. (2000). Chem. Rev. 100, 1025-1074.